

When predicting products...

1. Figure out the type of reaction that will most likely happen.
2. If it's combustion, the products will be $\text{CO}_2 + \text{H}_2\text{O}$!
3. Ions must combine so the total charge equals zero.
4. Cations are written first, and anions are written second in a compound.
5. Opposites attract, so cations pair with anions.
6. Don't forget what happens when a diatomic element is unbonded!

Question #1



Question #2



Question #3



- A. AgClMg
- B. $\text{Cl}_2\text{Mg} + \text{Ag}$
- C. $\text{MgAg} + \text{Cl}_2$
- D. $\text{MgCl} + \text{Ag}$
- E. $\text{MgCl}_2 + \text{Ag}$

Question #4



Question #5



- A. $\text{AgNO}_3\text{MgCl}_2$
- B. $\text{AgCl}_2 + \text{MgNO}_3$
- C. $\text{AgMg} + \text{NO}_3\text{Cl}$
- D. $\text{ClAg} + (\text{NO}_3)_2\text{Mg}$
- E. $\text{AgCl} + \text{Mg}(\text{NO}_3)_2$