

Topic: Stoichiometry	Name:	Date:
Questions/Main Ideas:	Notes:	
What do the coefficients in an equation represent?		
What is a mole to mole ratio?		
Why do you need stoichiometry?	<ul style="list-style-type: none"> To convert from amounts (moles, grams, particles) of one chemical to amounts of _____ To do this, we need the ratios from a _____ 	
What steps are needed to solve Stoichiometry problems	<ul style="list-style-type: none"> Balance your chemical equation. Write the given in the problem Convert to moles Write the mole to mole ratio (from the balanced chemical equation coefficients) 	
	<u>Moles of the material are you trying to get to</u> <u>Moles of the material are you leaving from</u>	
	<ul style="list-style-type: none"> Convert to what you are asked for 	
Convert moles to mass	Determine the mass of lithium hydroxide produced when 3.80 moles of lithium nitride reacts with water according to the following equation: $\text{Li}_3\text{N} + 3\text{H}_2\text{O} \rightarrow \text{NH}_3 + 3\text{LiOH}$	
Convert moles to moles	Determine the moles of lithium hydroxide produced when 3.80 moles of lithium nitride reacts with water according to the following equation: $\text{Li}_3\text{N} + 3\text{H}_2\text{O} \rightarrow \text{NH}_3 + 3\text{LiOH}$	
Convert mass to mass	How many grams of sodium chloride are produced from 53.5 g of sodium and excess amounts of chlorine gas?	
Independent problem	<i>Work this one on the back</i>	
Summary and Questions I have: (you must summarize what a mole ratio does and how to solve a stoichiometry problem)		

