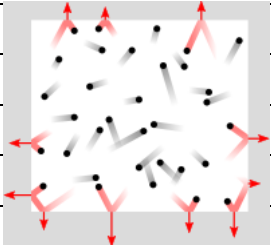


Topic: Kinetic Molecular Theory	Name:	Date:
Questions/Main Ideas:	Notes:	
What is kinetic molecular theory?	A model to	
What are the basics of kinetic molecular theory?	1. Gas molecules behave	
	2. Gas particles move	
	3. Particles are <i>really</i> small compared to	
What are some assumptions of kinetic molecular theory?	1. We assume that the volume of individual gas particles is so _____ we can ignore it	
	2. We assume that gas particles don't _____ each other	
	3. We assume that when gas particles collide, no _____ energy is lost	
What are some macroscopic properties of gases (things we can observe)?	1. Gases behave as _____ (they can flow, they can be poured)	
	2. Gases have low _____	
	3. Gases are easily _____	
	4. Gases completely fill a container and	
So, what is pressure ?	Gas pressure results when	
	<ul style="list-style-type: none"> • These collisions exert a force over an area (which is pressure) 	
	The amount of pressure depends on how _____ and how _____	
	the gas molecules hit the container	
	← In the picture, circle all the particles which are causing pressure .	
How do I convert between different units of pressure?	1.00 atm = 760. mm Hg = 760. torr = 101 kPa	
	<i>1 atmosphere (atm) is the usual pressure we feel under our atmosphere</i>	
The pressure on top of Mount Everest is 33.7 kPa. What is this pressure in atmospheres?		
<i>Show all work using D.A. →</i>		
Summary and Question(s) I have:		

