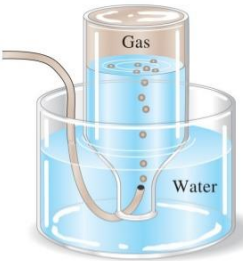
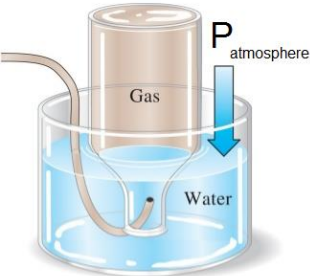


Topic: Dalton's Law	Name:	Date:
Questions/Main Ideas:	Notes:	
If I have a mixture of gases, what is the total pressure?	If I have a mixture of gases, what is the total pressure?	
		
	$P_{\text{total}} = P_{\text{gas}} + P_{\text{water vapor}}$	
	P_{air} = _____	
	How do you collect a gas that you generate in an experiment?	
	You can collect the gas _____	
	Gases are _____ dense than water, so they will rise to the top of a jar filled with water.	
	When you collect a gas over water, it has another gas mixed with it. It now has _____ mixed with the gas you were trying to collect. The _____ the temperature, the _____ the water vapor pressure.	
	How can you determine the total pressure?	
	If the water levels are the same on the _____ and _____, then total pressure (P_{total}) is equal to the _____ of the atmosphere (aka _____ pressure).	
		
	If the water levels are equal...	
	$P_{\text{total}} = P_{\text{atmosphere}}$	
	<u>Example:</u> If hydrogen gas is collected over water, and the barometric pressure is 1.013 atm, and the water vapor pressure is 7.5 mm Hg, what is the pressure of the hydrogen gas? (<i>Show work</i>)	
	Summary and Question(s) I have:	