

Topic: Ideal Gas Law	Name:	Date:
Questions/Main Ideas:	Notes:	
<i>Quick review:</i> List all of the gas variables we have used so far. Any important variable(s) missing?		
What is the Ideal Gas Law?	P =	n =
	V = (must be in ___)	T = (must be in ___)
What is the universal gas constant, R? How do I choose the correct one to use? (<i>answer</i> →)	R = ___ L·kPa/mol·K = ___ L·atm/mol·K = ___ L·mmHg/mol·K	
<u>Example 1:</u> What mass of ammonia is required to fill a 14.88 L bottle to a pressure of 199 kPa at 25°C?		
<u>Ex. #2:</u> What is the volume of 1.00 mole of a gas at STP?		
Why is it called the “ideal” gas law? What exactly is an ideal gas?	<i>An “ideal gas” follows those assumptions of Kinetic Molecular Theory:</i>	
	1.	(they actually do)
	2.	(they actually can)
	3.	(actually, gas particles lose a <u>tiny</u> amount of energy as friction when they collide)
When do these assumptions start to become significant in our ideal gas law calculations?		
Summary and Question(s) I have:		