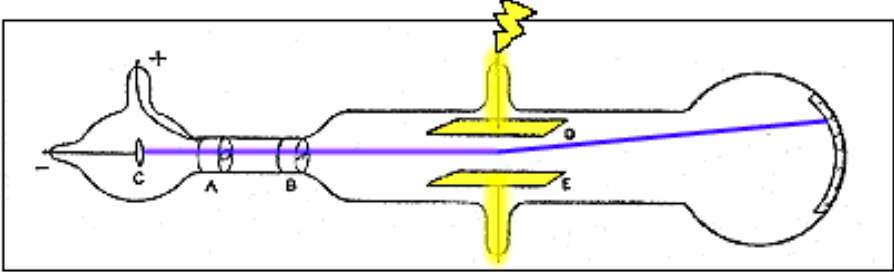


Topic: Thomson	Name:	Date:
<b>Questions/Main Ideas:</b>	<b>Notes:</b>	
Before Thomson's experiment, what was the scientific consensus?	Atoms are supposed to be the _____ unit of matter. Atoms are electrically _____ (that is, they have _____ charge).	
How did Thomson set up his cathode ray tube experiment?	Thomson took an empty	
	When connected, a beam	
	This <i>cathode ray</i> would glow _____ when it struck a special coating	
What was so surprising about what Thomson observed?	The beam was _____ (moved) by electric or magnetic fields. Only particles with _____ are affected by electric/magnetic fields. The atom is supposedly the _____ unit of matter, with no charge. So, this beam must be made up of _____, not atoms.	
Label the <b>cathode</b> , <b>cathode ray</b> , <b>positive plate</b> , and <b>negative plate</b> .		
So, <b>what</b> is this beam made of?	The beam was <b>attracted</b> to the _____ charged plate, and <b>deflected</b> from the _____ charged plate. Since opposite charges _____, that means the beam must be _____ charged. What is it made of? <del>Corpuscles</del> _____	
What else did Thomson determine about this particle?	He could not determine the particle's mass, or its charge. However, by measuring the _____ of the beam at the end of his tube, and applying a _____ and _____ field <i>at the same time</i> , he was able to determine how the charge and mass were <b>related</b> .	
<b>Where</b> are these electrons inside the atom? ( <i>draw the "Plum Pudding" model below</i> )	Thomson <u>assumed</u> the electrons were _____ throughout the atom like chocolate chips in a cookie (this is called the "_____ Pudding" model) Since the atom was _____ overall, he <u>assumed</u> the entire atom was _____ (to balance out the negative _____)	
<b>Summary and Question(s) I have:</b> ( <i>be sure to explain Thomson's evidence</i> )		
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