

Topic: Electronegativity and IE	Name:	Date:
Questions/Main Ideas:	Notes:	
Review: What are the trends for effective nuclear charge?	Across a period, effective nuclear charge _____ because	
	Down a group, effective nuclear charge is _____.	
What is ionization energy?	the energy	
	<i>(The energy you have to "pay" to take an electron away!)</i>	
	The more positive the number, the	
	Ionization energy _____ going down a group because	
	Ionization energy _____ going left to right across a period because	
Are there any exceptions to the trend for ionization energy?	The noble gases have extremely _____ ionization energies because	
Which of these elements has the highest <u>ionization energy</u> ?	A. Oxygen B. Francium C. Chromium D. Calcium E. Sulfur <i>Explain why:</i>	
What is electronegativity?	the ability of	
	The more positive the number, the	
	<i>Think of electronegativity as a game of "tug of war" for the electrons in a covalent bond – the element with the higher electronegativity will have an unfair share of the electrons in the covalent bond</i>	
	Electronegativity _____ going down a group because of	
	Electronegativity _____ left to right across a period because of	
Are there any exceptions to the trend for electronegativity?	Fluorine is the _____ electronegative element because of a combination of its _____ effective nuclear charge and is _____ atomic size. The electrons are being pulled on very strongly.	
	The noble gases have _____ electronegativity because	
Which of these elements has the lowest <u>electronegativity</u> ?	A. Zinc B. Lithium C. Strontium D. Fluorine E. Boron <i>Explain why:</i>	
Summary and Question(s) I have: (be sure so summarize the meaning and trends for IE and electroneg):		

